Name:_____

Instructions: Fill-in the following table by providing the missing entry. Place "NA" if not applicable. *Check your answers after you complete several lines.*
Note: If the compound type is "covalent" include "N/A" in the Ions present column.

Chemical Name	**Compound Type**	**Ions present (if ionic)**	**Chemical Formula**
Ex. aluminum sulfide	ionic	[2 Al^{+}, 3 S^{2-}]	Al_2S_3
1. _____	_____	[Cr^{3+} , O^{2-}]	_____
2. zinc chloride	_____	_____	_____
3. _____	_____	_____	CO_2
4. _____	_____	_____	$Fe(CN)_2$
5. _____	_____	[Al^{3+} , SO_4^{2-}]	_____
6. arsenic trihydride	_____	_____	_____
7. ammonium chromate	_____	_____	_____
8. _____	_____	[Ca^{2+} , PO_4^{3-}]	_____
9. silicon tetrafluoride	_____	_____	_____
10._____	_____	_____	$CuCl_2$
11. iron(II) permanganate	_____	_____	_____
12._____	_____	[Co^{2+} , HPO_4^{2-}]	_____
13. potassium dichromate	_____	_____	_____
14. magnesium phosphide	_____	_____	_____
15._____	_____	[Na^{+} , CO_3^{2-}]	_____
16._____	_____	_____	$Ni(ClO_4)_2$
17._____	_____	[Cu^{2+} , $C_2H_3O_2^{-}$]	_____
18. cobalt(III) hydroxide	_____	_____	_____
19._____	_____	_____	$Cd(ClO_3)_2$

Chemical Name	Compound Type	Ions present (if ionic)	Chemical Formula
20. _____	_____	[Mn^{2+} , BrO^{-}]	_____
21. titanium(IV) oxalate	_____	_____	_____
22. calcium hypoiodite	_____	_____	_____
23. mercury(I) chloride	_____	_____	_____
24. _____	_____	_____	N_2O_4
25. chlorine dioxide	_____	_____	_____
26. _____	_____	[Ag^{+} , NO_3^{-}]	_____
27. _____	_____	[Ni^{2+} , OH^{-}]	_____
28. sodium bicarbonate*	_____	_____	_____
29. *other name: _____			
30. tin(IV) chloride	_____	_____	_____
31. mercury(II) nitrate	_____	_____	_____
32. _____	_____	[Ga^{3+}, S^{2-}]	_____
33. _____	_____	[Pb^{4+}, O^{2-}]	_____
34. hydrogen chloride	_____	_____	_____
35. zinc acetate	_____	_____	_____
36. _____*	_____	_____	$FeCl_3$
*other name: _____			
37. calcium bisulfate	_____	_____	_____
other name: _____			
38. aluminum bisulfite*	_____	_____	_____
*other name: _____			

Chemical Name	Compound Type	Ions present (if ionic)	Chemical Formula
39._____*	_____	_____	$Al(HSO_4)_3$
*other name:_____			
40._____	_____	_____	SeO_3
41._____	_____	_____	BCl_3
42. cuprous nitrate	_____	_____	_____
43._____	_____	[Hg_2^{2+} , Br^-]	_____
44. nitric acid	_____	_____	_____
45. lead(II) acetate	_____	_____	_____
46. sulfurous acid	_____	_____	_____
47. chromic acid	_____	_____	_____
48._____	_____	[H^+, Cl^-]	_____
49. mercurous sulfide	_____	_____	_____
50. *sodium hydrogen sulfate	_____	_____	_____
*other name: _____			
51._____	_____	_____	HNO_2
52._____	_____	_____	$HBrO_4$
53. sulfur hexafluoride	_____	_____	_____
54._____	_____	_____	HBr(aq)